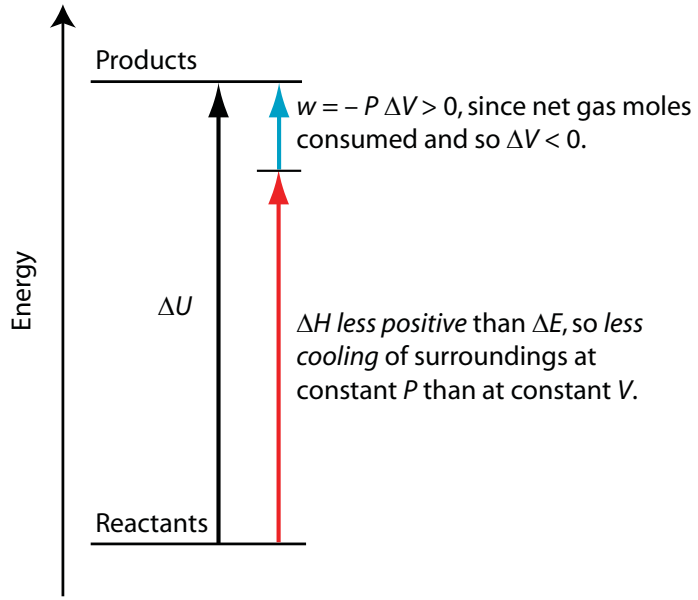
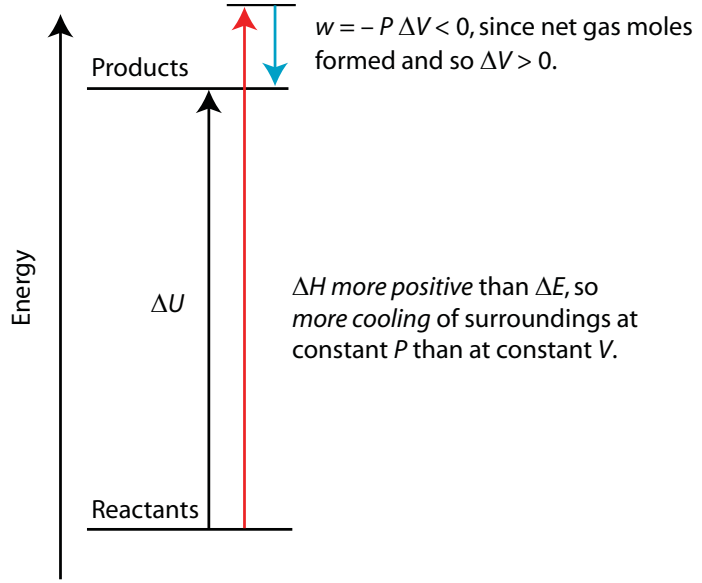
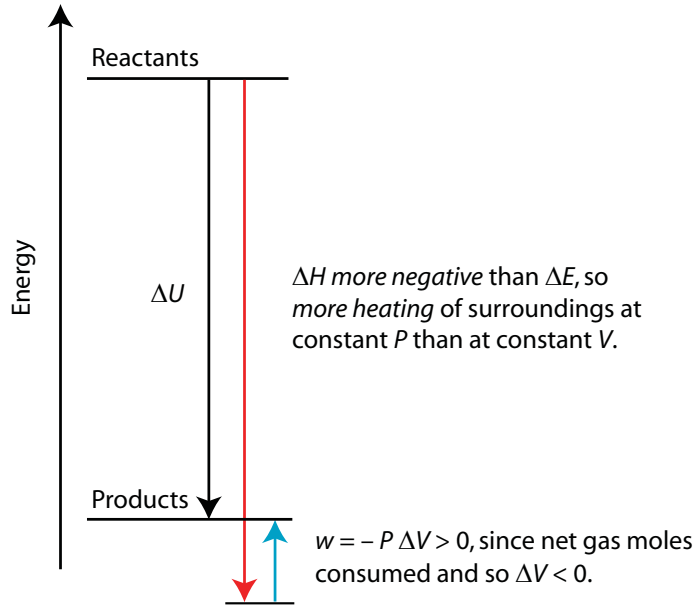
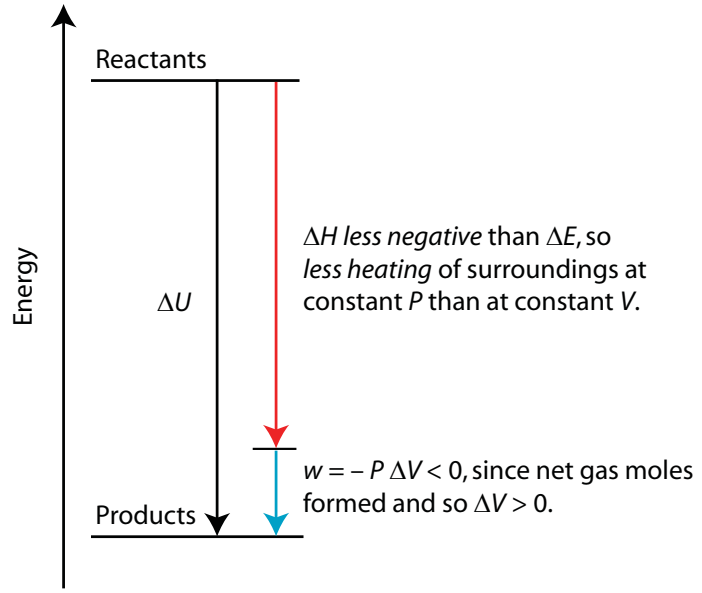
Relationship between energy change, ΔU (black path), and enthalpy change, ΔH (red path), according to (1) whether products have more (upper figures) or less (lower figures) energy than reactants, and (2) whether work (cyan path) is done on the system (left figures) or on the surrounding (right figures). When no work is done at constant P, then ΔU and ΔH have the same value.









Relationship between energy change, ΔU (black path), and enthalpy change, ΔH (red path), according to (1) whether products have more (upper figures) or less (lower figures) energy than reactants, and (2) whether work (cyan path) is done on the system (left figures) or on the surrounding (right figures). When no work is done at constant P, then ΔU and ΔH have the same value.