

- 1 0.001 mol of a strong acid H-A is placed in 1.00 L water at 25°C. Before equilibrium is established, what is (H-A)?
- A 0
 - B 0.001
 - C 10^{-7}
 - D More information is needed.
- 2 0.001 mol of a strong acid H-A is placed in 1.00 L water at 25°C. Before equilibrium is established, what is (H₃O⁺)?
- A 0
 - B 0.001
 - C 10^{-7}
 - D More information is needed.
- 3 0.001 mol of a strong acid H-A is placed in 1.00 L water at 25°C. Before equilibrium is established, what is (:OH⁻)?
- A 0
 - B 0.001
 - C 10^{-7}
 - D More information is needed.
- 4 0.001 mol of a strong acid H-A is placed in 1.00 L water at 25°C. After equilibrium is established, what is (H₃O⁺)?
- A ~0
 - B ~0.001
 - C ~ 10^{-7}
 - D More information is needed.
- 5 0.001 mol of a strong acid H-A is placed in 1.00 L water at 25°C. After equilibrium is established, what is (H-A)?
- A Exactly 0
 - B ~0.001
 - C ~ 10^{-7}
 - D More information is needed.