1 The expression for the equilibrium constant for solubility equilibrium M2X(s) <--> 2 M+(aq) + X2-(aq) is ...

A Ksp = (2 M+)(X2-)/(M2X)
B Ksp = (2 M+)2(X2-)/(M2X)
C Ksp = (2 M+)2(X2-)
D Ksp = (M+)2(X2-)

2 The equilibrium constant for M2X(s) <--> 2 M+(aq) + X2-(aq) is Ksp = (M+)2(X2-) = 8 x 10-11. Assume a maximum of s moles of M2X(s) can dissolve. What is true for s?

A Ksp = (2 s)(s)
B Ksp = (2 s)2(s)
C Ksp = (s)2(s)
D None of the above

3 Ksp for M2X(s) <--> 2 M+(aq) + X2-(aq) is 8 x 10-11. Assume a maximum of s moles of M2X(s) can dissolve. What is true for s if M+(aq) is initially 0.1 M?

A Ksp = (2*0.1)(s)
B Ksp = (2*0.1)2(s)
C Ksp = (0.1)2(s)
D None of the above