1 What is the expression for the equilibrium constant of an aqueous solution of the oxoacid conjugate base :OA-?

A Kb = (:OA-)(:OH-)
B Kb = (:OH-)(:OA-)/(H-OA)
C Kb =(H-OA)(:OH-)/(:OA-)
D Kb = (H3O+)(:OA-)/(HOA)

2 H-OA is a weak acid with Ka = 10-5. What is the pH of 0.1 M conjugate base :OA-?

A < 7
B 7
C > 7
D More information needed

3 H-OA is a weak acid with Ka = 10-5 at 25oC. What is K for
H-OH + :OA- <=> H-OA + :OH- at 25oC?

A K > 10-5
B K = 10-5
C K < 10-5
D More information needed

4 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25oC. Before equilibrium is established, what is (OA-)?

A 0
B 0.001
C 10-7
D More information needed

5 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25oC. Before equilibrium is established, what is (HOA)?

A 0
B 0.001
C 10-7
D More information needed

6 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25oC. Before equilibrium is established, what is (OH-)?

A 0
B 0.001
C 10-7
D More information needed

7 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25oC. After equilibrium is established, what is (:OA-)?

A 0
B ~0.001
C 10-7
D More information needed

8 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25oC. After equilibrium is established, what is (HOA)?

A 0
B 0.001
C 10-7
D More information needed

9 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25oC. After equilibrium is established, what is (OH-)?

A 0
B 0.001
C 10-7
D More information needed