1 There are alternative definitions of acid-base reactions. In the BRONSTED-LOWRY definition, ...

A A base contributes a nonbonding pair of electrons (:) to form a bond
B An acid accepts a proton (H+)
C A base accepts a nonbonding pair of electrons (:) to form a bond
D None of the above

2 When a BRONSTED-LOWRY (BL) acid reacts, a bond between a proton (H+) and the rest of the molecule or ion is broken. A LEWIS acid ...

A forms a bond with a proton (H+)
B also always breaks a bond with proton (H+)
C always donates a nonbonding pair of electrons (:) to form a bond with a proton (H+)
D None of the above

3 A BL acid always donates a proton (H+). A LEWIS (L) acid is the result of a bond with a nonbonding pair of electrons. Which of the following is true?

A For both a BL base reaction and an L base reaction with an acid, a bond is formed.
B For both a BL acid reaction and an L acid reacting with a base, a proton bond is always broken.
C Neither of the above
D Both of the above

4 For both a BL base and an L base reacting with an acid, a bond is formed. This means the relative strength of an acid must due to ...

A pH
B pressure
C temperature
D bond strength
E All of the above