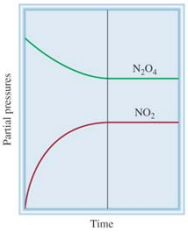
Reaction quotient versus time

General Chemistry, CH102 Spring 2011

1. The figure shows how the partial pressures of the N_2O_4 and NO_2 **change with time** due to the chemical reaction $N_2O_4 \rightarrow 2\ NO_2$ for **certain initial conditions**. At these initial conditions, the following is known about the chemical reaction.



0% 1. It is **spontaneous**
0% 2. It is **at equilibrium**
0% 3. It is **non-spontaneous**
0% 4. Its spontaneity is **not known** without further information

0 of 5

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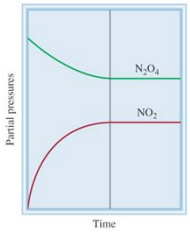
2. The figure shows how the partial pressures of the N_2O_4 and NO_2 **change with time** due to the chemical reaction $N_2O_4 \rightarrow 2\ NO_2$ for **certain initial conditions**. At these initial conditions, the following is known about the ratio ***Q/K***.



0% 1. It is **greater** than 1
0% 2. It is **equal** to 1
0% 3. It is **less** than 1
0% 4. The ratio is not known without further information

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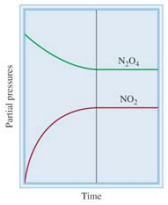
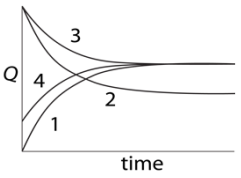
3. For the reaction $N_2O_4 \rightarrow 2\ NO_2$ and **certain initial conditions**, the figure on the left shows how the partial pressures of the N_2O_4 and NO_2 change with time. Which curve on the right shows the corresponding **change of Q with time**?





0% 1. 1
0% 2. 2
0% 3. 3
0% 4. 4
0% 5. None

0 of 5

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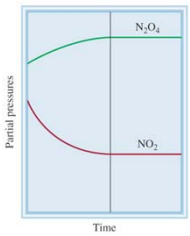
4. The figure shows how the partial pressures of the N_2O_4 and NO_2 **change with time** due to the chemical reaction $N_2O_4 \rightarrow 2\ NO_2$ for **certain initial conditions**. At these initial conditions, the following is known about the chemical reaction.



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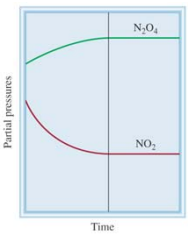

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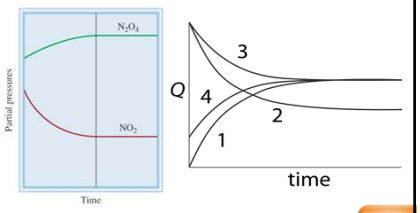
6. For the reaction $N_2O_4 \rightarrow 2\ NO_2$ and **certain initial conditions**, the figure on the left shows how the partial pressures of the N_2O_4 and NO_2 change with time. Which curve on the right shows the corresponding **change of Q with time**?



0% 1. 1
0% 2. 2
0% 3. 3
0% 4. 4
0% 5. None

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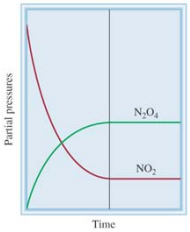
7. The figure shows how the partial pressures of the N_2O_4 and NO_2 **change with time** due to the chemical reaction $N_2O_4 \rightarrow 2\ NO_2$ for **certain initial conditions**. At these initial conditions, the following is known about the chemical reaction.



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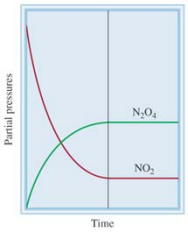

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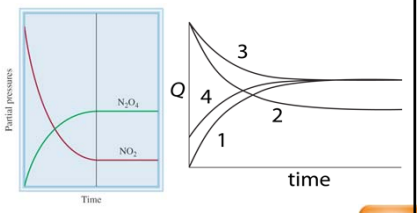
9. For the reaction $N_2O_4 \rightarrow 2\ NO_2$ and certain initial conditions, the figure on the left shows how the partial pressures of the N_2O_4 and NO_2 change with time. Which curve on the right shows the corresponding change of Q with time?



0% 1. 1
0% 2. 2
0% 3. 3
0% 4. 4
0% 5. None

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