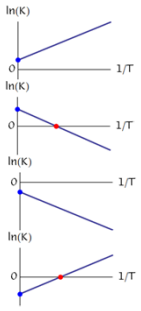
Effect of temperature on equilibrium

General Chemistry, CH102 Spring 2011

1. For **exothermic** reactions in which **entropy increases**, which figure (1, 2, 3 or 4, top to bottom) corresponds to ln(*K*) versus 1/*T*?

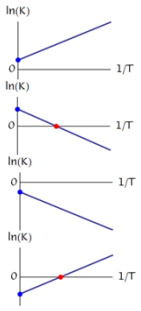


96% 1. Figure 1
3% 2. Figure 2
0% 3. Figure 3
1% 4. Figure 4

2. For **exothermic** reactions in which **entropy increases**, what temperature range will result in the **greatest amount of products**?

63% 1. Very low *T*
23% 2. Very high *T*
14% 3. The amount will be the same at all *T*
0% 4. More information needed

3. For **endothermic** reactions in which **entropy increases**, which figure (1, 2, 3 or 4, top to bottom) correspond to ln(*K*) versus 1/*T*?



0% 1. Figure 1
0% 2. Figure 2
0% 3. Figure 3
0% 4. Figure 4

0 of 5

10 Countdown Timer On Slide

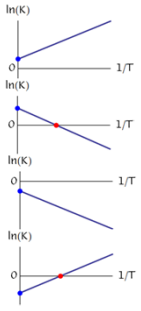
4. For **endothermic** reactions in which **entropy increases**, what temperature range will result in the **greatest amount of products**?

0% 1. Very low *T*
0% 2. Very high *T*
0% 3. The amount will be the same at all *T*
0% 4. More information needed

0 of 5

10 Countdown Timer On Slide

5. For **exothermic** reactions in which **entropy decreases**, which figure (1, 2, 3 or 4, top to bottom) correspond to ln(*K*) versus 1/*T*?



0% 1. Figure 1
0% 2. Figure 2
0% 3. Figure 3
0% 4. Figure 4

0 of 5

10 Countdown Timer On Slide

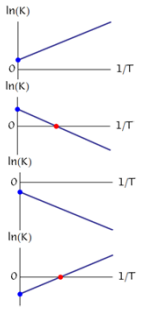
6. For **exothermic** reactions in which **entropy decreases**, what temperature range will result in the **greatest amount of products**?

0% 1. Very low *T*
0% 2. Very high *T*
0% 3. The amount will be the same at all *T*
0% 4. More information needed

0 of 5

10 Countdown Timer On Slide

7. For **endothermic** reactions in which **entropy decreases**, which figure (1, 2, 3 or 4, top to bottom) correspond to ln(K) versus 1/T?



0% 1. Figure 1
0% 2. Figure 2
0% 3. Figure 3
0% 4. Figure 4

0 of 5

10 Countdown Timer On Slide

8. For **endothermic** reactions in which **entropy decreases**, what temperature range will result in the **greatest amount of products**?

0% 1. Very low *T*
0% 2. Very high *T*
0% 3. The amount will be the same at all *T*
0% 4. More information needed

0 of 5

10 Countdown Timer On Slide

9. For the reaction $2\ H_2O(g) \rightarrow 2\ H_2(g) + O_2(g)$, what temperature range will result in the **greatest amount of products**?

0% 1. Very low *T*
0% 2. Very high *T*
0% 3. The amount will be the same at all *T*
0% 4. More information needed

0 of 5

10 Countdown Timer On Slide

10. The reaction $N_2(g) + 2\ H_2(g) \rightarrow 2\ N_2H_4(l)$ is endothermic. What temperature range will result in the **greatest amount of products**?

0% 1. Very low *T*
0% 2. Very high *T*
0% 3. The amount will be the same at all *T*
0% 4. More information needed

0 of 5

10 Countdown Timer On Slide

11. The reaction $N_2O_4(g) \rightarrow 2\ NO_2(g)$ is exothermic. What temperature range will result in the **greatest amount of products**?

0% 1. Very low *T*
0% 2. Very high *T*
0% 3. The amount will be the same at all *T*
0% 4. More information needed

0 of 5

10 Countdown Timer On Slide