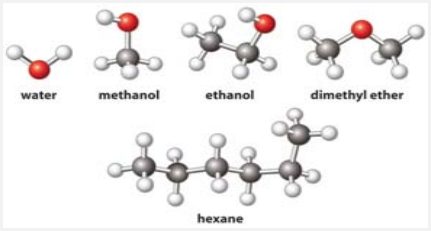


[TP] Identify the compound with the smallest energy of vaporization (kJ/mol).

25% 1. water
25% 2. ethanol
25% 3. dimethyl ether
25% 4. hexane



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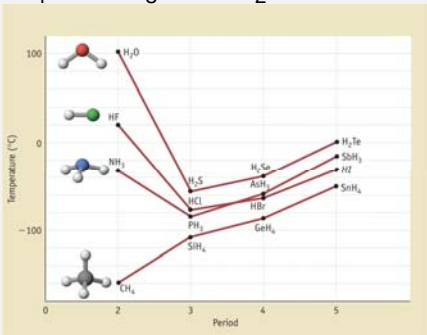
Lecture 16 CH101 A1 (MWF 9 am)

Friday, October 14, 2016

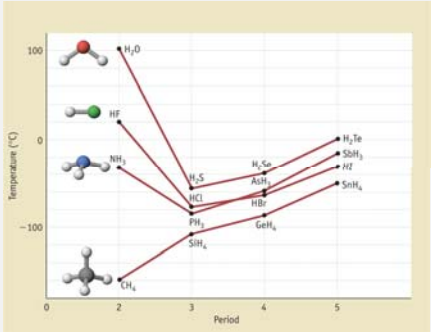
- Complete: Relative boiling points

Next lecture: Review: Intermolecular forces; dissolving ionic solids, solubility rules (memorize solubility guidelines fig 6.28 p 181); precipitation reactions; concentrations after precipitation; ionization of molecular solutes; self-ionization of water

Why is $SnH_4 < SbH_3 < TeH_2$?



Why is $CH_4 < SiH_4 < GeH_4 < SnH_4$?



[TP] The correct order of polarity of HBr, HCl and HI is (least to most) ...

17% 1. HBr < HCl < HI
17% 2. HBr < HI < HCl
17% 3. HCl < HBr < HI
17% 4. HCl < HI < HBr
17% 5. HI < HCl < HBr
17% 6. HI < HBr < HCl

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[TP] The correct order of boiling point of HBr, HCl and HI is (lowest to highest) ...

17% 1. HBr < HCl < HI
17% 2. HBr < HI < HCl
17% 3. HCl < HBr < HI
17% 4. HCl < HI < HBr
17% 5. HI < HCl < HBr
17% 6. HI < HBr < HCl

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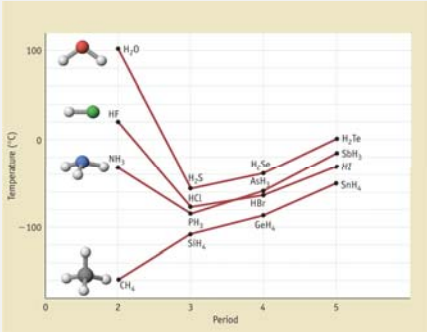
Why is HCl < HBr < HI?



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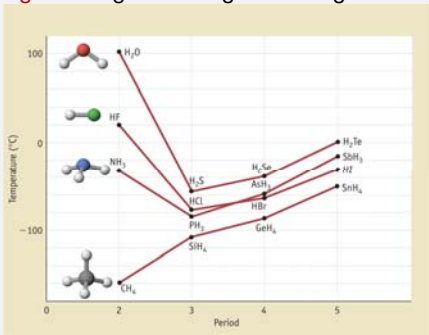
Why is $NH_3 > PH_3 < AsH_3 < SbH_3$?



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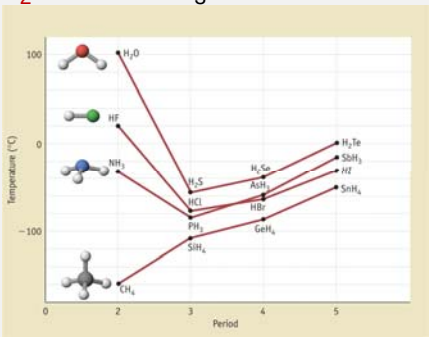
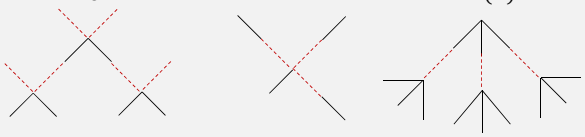
Why is OH_2 > FH > NH_3?

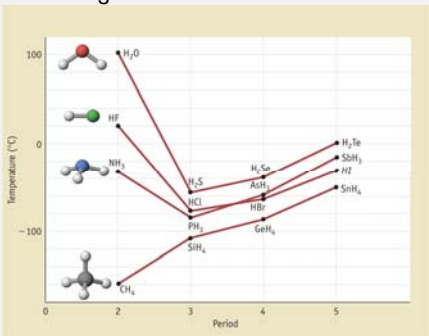


Why is OH_2 > FH > NH_3?

H-O-H H-F H-N(H)-H



Why is FH > NH_3?



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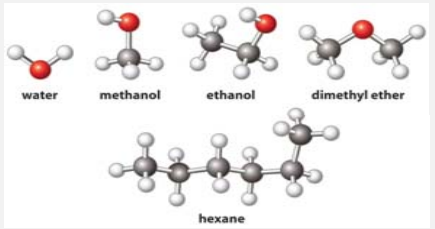

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