Lecture 13 (10/7/20)

- Reading: Ch6; 187-189, 204-205, 218-219
- Problems: Ch6 (text); 2, 3, 5, 6
  Ch6 (study guide); *1, 22* (facts)

NEXT (after Exam 2)

- Reading: Ch6; 190-191, 194-195, 197-198
- Problems: Ch6 (text); 7, 24
  Ch6 (study guide); *13, 4* (facts)

Lecture 13 (10/7/20)

OUTLINE

ENZYMES: Binding & Catalysis

A. Binding
   1. Binding curves; **How tight?**
      a. Hyperbolic –saturation
      b. Sigmoidal –cooperativity

B. Catalysis
   1. Catalytic power
      a. Proficiency
      b. assay of rate
      c. rate *versus* [E]

C. Nomenclature
   1. Reaction Nomenclature (kinetic mechanism)
   2. Enzyme helpers: Cofactors
   3. Enzyme Nomenclature (names)
      a. Trivial
      b. Enzyme Commission (EC#)

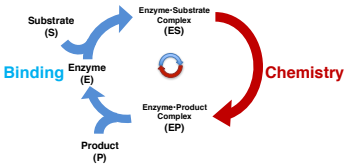
D. Catalysis (reprise)
   1. Transition State Theory
      a. Energetics (thermodynamics) vs. kinetics
      b. Lower activation energy; negative $\Delta\Delta G^{\ddagger}$
   2. Catalytic Strategies

Enzymes

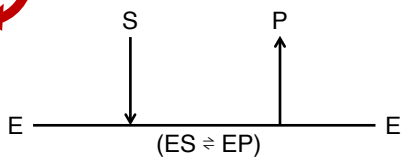
4) Enzyme nomenclature
   a. Reaction Nomenclature
   b. Enzyme helpers (cofactors)
   c. Naming enzymes

Reaction Nomenclature

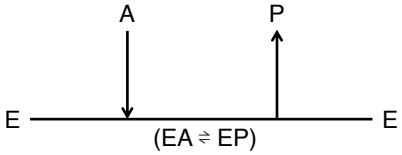
Enzymes



$E + S \rightleftharpoons ES \rightleftharpoons EP \rightleftharpoons E + P$



- Reaction coordinate written as a horizontal line
- Substrates binding with arrow down; substrates are denoted in order A, B, C
- Products released with arrow up; products are denoted in order P, Q, R
- Enzyme forms are denoted in order as E, F, G
- Define the number of substrates and/or products separately as *uni*=1, *bi*=2, *ter*=3
- Define the order of binding and release separately as *ordered* or *random*
- Define the relationship of substrates-to-products as *sequential* or *ping-pong*



What about more complicated reactions?

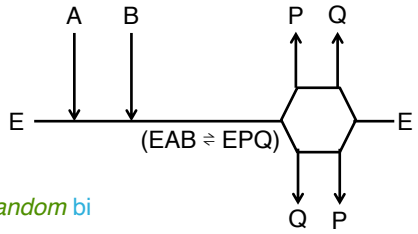
This is a uni-uni reaction.
(No need to designate last two when there is only one substrate and one product)

Reaction Nomenclature Enzymes

(a) Enzyme reaction involving a ternary complex

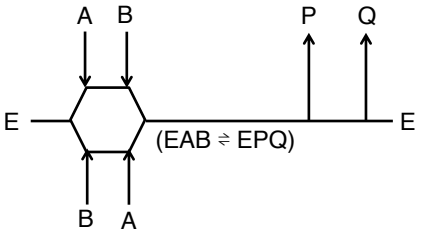
Ordered

$E + S_1 \rightleftharpoons ES_1 \overset{S_2}{\rightleftharpoons} ES_1S_2 \longrightarrow E + P_1 + P_2$



Sequential ordered bi *random* bi

Random order



Sequential random bi *ordered* bi

Reaction Nomenclature Enzymes

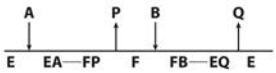
Cleland nomenclature

Two-substrate/Two-product reactions: bi bi

Ordered bi bi

Sequential ordered bi *ordered* bi
Sequential ordered bi bi
Ordered bi bi

Sequential random bi *random* bi
Sequential random bi bi
Random bi bi



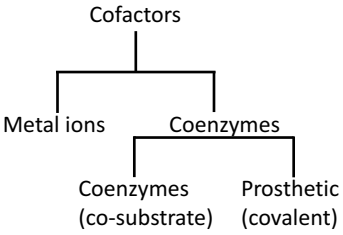
Ping-pong ordered bi *ordered* bi
Ping-pong bi bi

Which is the bi bi reaction catalyzed by LDH?

Enzyme Helpers

Enzymes

Cofactors are small molecules that some enzymes require for activity. The two main classes of cofactors are coenzymes (organic molecules derived from vitamins) and metals. Covalently bound coenzymes are called prosthetic groups.



An enzyme with its cofactor is a **holoenzyme**. Without the cofactor, the enzyme is called an **apoenzyme**.

Cofactors	Enzyme
Coenzyme†	
Thiamine pyrophosphate (TPP)	Pyruvate dehydrogenase
Flavin adenine dinucleotide (FAD)	Monoamine oxidase
Nicotinamide adenine dinucleotide (NAD^+)	Lactate dehydrogenase
Pyridoxal phosphate (PLP)	Glycogen phosphorylase
Coenzyme A (CoA)	Acetyl CoA carboxylase
Biotin	Pyruvate carboxylase
6′-Deoxyadenosyl cobalamin	Methylmalonyl mutase
Tetrahydrofolate	Thymidylate synthase
Metal	
Zn^{2+}	Carbonic anhydrase
Mg^{2+}	*Eco*RV
Ni^{2+}	Urease
Mo	Nitrogenase
Se	Glutathione peroxidase
$Mn^{2+\leftrightarrow 3+}$	Superoxide dismutase
K^+	Acetoacetyl CoA thiolase

*The enzymes listed are examples of enzymes that employ the indicated cofactor.
†Often derived from vitamins, coenzymes can be either tightly or loosely bound to the enzyme.

Naming Enzymes

Enzymes

***Trivial:**

Name	Reaction Catalyzed

- Nearly all enzymes end with the suffix of "-ase."
- Generally, the names are of the form "substrate or product – reaction catalyzed." For example, lactate dehydrogenase is for an enzyme that removes a hydrogen (plus $2e^-$, i.e., a hydride) from lactate, yielding the carbonyl in pyruvate.
- There are two ways of naming enzymes; 1) Trivial and 2) Systematic

* Not all possible types listed

• Bullets are those also given as systematic

Naming Enzymes

Enzymes

*Trivial:

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- Generally, the names are of the form "substrate or product – reaction catalyzed." For example, lactate dehydrogenase is for an enzyme that removes a hydrogen (plus $2e^-$, i.e., a hydride) from lactate, yielding the carbonyl in pyruvate.
- There are two ways of naming enzymes; 1) Trivial and 2) Systematic

* Not all possible types listed

• Bullets are those also given as systematic

Name	Reaction Catalyzed
-dehydrogenase	redox/hydride transfer
•	*lactate dehydrogenase*
•	*glyceraldehyde-3-phosphate dehydrogenase*
-oxidase	redox/O_2 as oxidizer
	cytochrome oxidase
•	*glucose oxidase*
-oxygenase	redox/O_2 incorporated
	cyclooxygenase
	Ribulose Bisphosphate Carboxylase Oxygenase
-hydroxylase	redox/-OH incorporated
	tyrosine hydroxylase
•	*phenylalanine hydroxylase*
-kinase	transfer/P_i into substrate from ATP
•	*hexose kinase*
	protein kinase A
-hydrolase	hydrolysis with H_2O
• (esterase, deacylase)	*trypsin*
•	*phospholipase C*
-phosphorylase	hydrolysis with P_i instead of H_2O
	glycogen phosphorylase b
	Thymidine phosphorylase
-mutase	move P_i from one part of molecule to another
	phosphoglycerate mutase
•	*phosphoglucose mutase*
-isomerase	configuration change
•	*triosephosphate isomerase*
•	*phosphogluco isomerase*
-synthase	synthesis
•	*fatty acid synthase*
	nitric oxide synthase
-synthetase	synthesis that requires ATP
•	*aminoacyl-tRNA synthetases*
	acyl-CoA synthetase

Great website: EC numbers
(https://www.qmul.ac.uk/sbcs/iubmb/enzyme/)

Systematic

1.1.1.1

Type Sub-type Sub-class enzyme specific*

Types	Reaction	Examples

Great website: EC numbers
(http://www.chem.qmul.ac.uk/iubmb/enzyme/)

Systematic

1.1.1.1
Type Sub-type Sub-class enzyme specific*

Types	Reaction	Examples	EC#	Name	Trivial Name
-oxidoreductase	redox	At alcohol → Uses NAD	1.1.1.28	L-lactate:NADH oxidoreductase	lactate dehydrogenase
			1.2.1.59	D-glyceraldehyde-3-phosphate:NAD+ oxidoreductase	glyceraldehyde-3-phosphate dehydrogenase
			1.1.3.4	β-D-glucose:oxygen 1-oxidoreductase	glucose oxidase
			1.14.16.1	L-phenylalanine,tetrahydrobiopterin:oxygen oxidoreductase (4-hydroxylating)	phenylalanine hydroxylase
-transferase	transfer	Nitrogen-group transfer → Primary amine	2.6.1.1	L-aspartate:2-oxoglutarate aminotransferase	Aspartate aminotransferase
			2.3.1.85	Acyl-CoA:malonyl-CoA C-acyltransferase	fatty acid synthase
			2.7.1.1	ATP:D-hexose 6-phosphotransferase	hexose kinase
-hydrolase	hydrolysis	Peptide bonds → Serine mechanism	3.4.21.4	*Serine endopeptidylamino acid hydrolase*	*trypsin*
			3.1.4.3	*Phosphatidylcholine:cholinephosphohydrolase*	*phospholipase C*
-lyase	bond cleavage	C-C bond cleavage → Aldehyde product	4.1.2.13	*Fructose 1,6-bisphosphate:triosephosphate lyase*	*aldolase*
			4.2.1.2	*(S)-malate hydro-lyase*	*fumarase*
			4.1.1.1	*2-oxo-acid carboxy-lyase (aldehyde-forming)*	*pyruvate decarboxylase*
-isomerase	configuration change	Intramolecular configuration change → Aldose:ketose	5.3.1.9	*Glucose-6-phophate isomerase*	*phosphoglucoisomerase*
			5.3.1.1	*D-glyceraldehyde-3-phosphate:aldose-ketose-isomerase*	*triosephosphate isomerase*
			5.1.1.4	*Proline racemase*	*proline racemase*
			5.4.2.11	*D-phosphoglycerate 2,3-phosphomutase (2,3-diphosphoglycerate-dependent)*	*phosphoglycerate mutase*
-ligase	synthesis	Ester linkage → Only 1 subclass	6.1.1.1	*Tyrosine aminoacyl-tRNA ligase*	*Tyr-aminoacyl-tRNA synthetases*
			6.4.1.1	*Pyruvate:carbon-dioxide ligase (ADP-forming)*	*pyruvate carboxylase*

*this is specific for alcohol dehydrogenase!

Naming Enzymes

Enzymes

Correlation of trivial and systematic:

OTHLIL

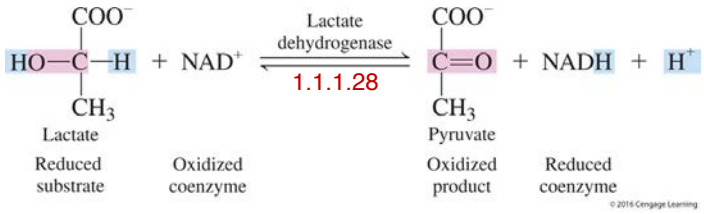
Main Classes	Selected Trivial Names	Type of Reaction Catalyzed
oxidoreductases	oxidases reductases dehydrogenases	oxidation of a substrate reduction of a substrate introduction of double bond (oxidation) by formal removal of two H atoms from a substrate, with one H being accepted by a coenzyme
transferases	transaminases kinases	transfer of an amino group between substrates transfer of a phosphate group between substrates
hydrolases	lipases proteases nucleases carbohydrases phosphatases	hydrolysis of ester linkages in lipids hydrolysis of amide linkages in proteins hydrolysis of sugar–phosphate ester bonds in nucleic acids hydrolysis of glycosidic bonds in carbohydrates hydrolysis of phosphate–ester bonds
lyases	dehydratases decarboxylases deaminases hydratases	removal of H_2O from a substrate removal of CO_2 from a substrate removal of NH_3 from a substrate addition of H_2O to a substrate
isomerases	racemases mutases	conversion of D isomer to L isomer, or vice versa transfer of a functional group from one position to another in the same molecule
ligases	synthetases carboxylases	formation of a new bond between two substrates, with participation of ATP formation of a new bond between a substrate and CO_2, with participation of ATP

Naming Enzymes

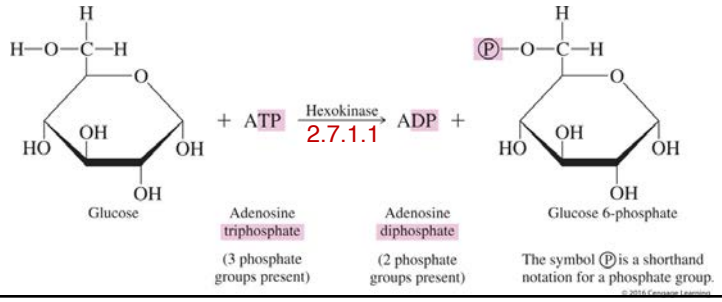
Enzymes

OTHLIL

Oxidoreductase: oxidation-reduction reaction



Transferase: transfer of functional group between molecules

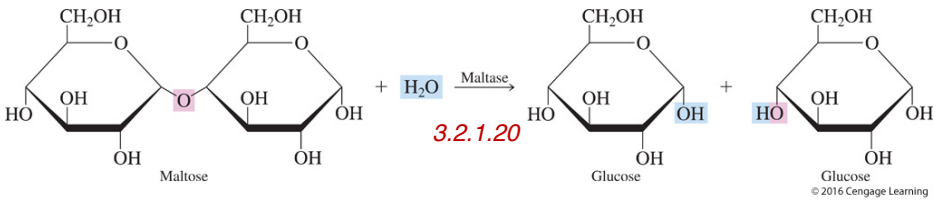


Naming Enzymes

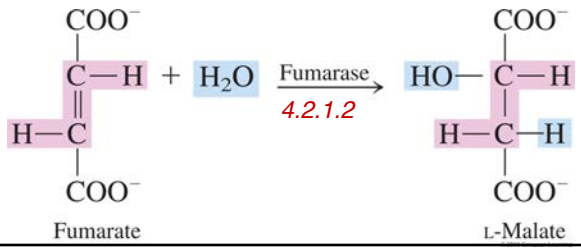
Enzymes

OTHLIL

Hydrolase: hydrolysis (addition of water breaks bond)



Lyase: addition of group to a double bond or
removal of group to form double bond without hydrolysis or oxidation

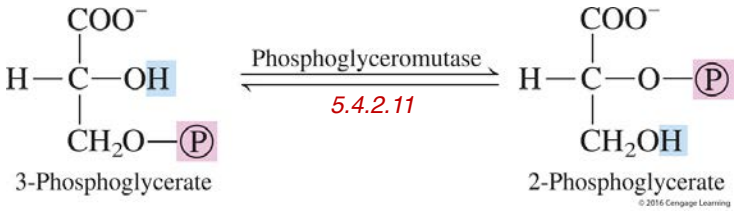


Naming Enzymes

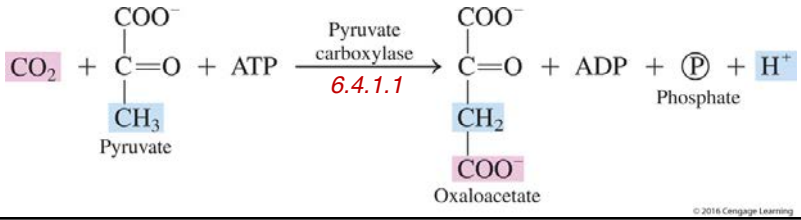
Enzymes

OTHLIL

Isomerase: transfer of functional group *within* a molecule (rearrangement)



Ligase: joining of two molecules (bond formation) coupled with ATP hydrolysis



Enzymes

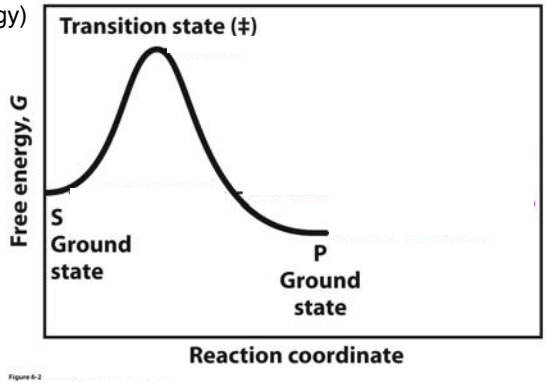
How do enzymes achieve these HUGE rate enhancements?

Let's review what governs the rates of chemical reactions: the transition state energy

OK, what does a similar plot look like using an enzyme?

Reaction Coordinate Diagram

Reaction rate $\propto \Delta G^{\ddagger}$ (activation energy)



⊕ $\Delta G^{\ddagger}$ – Enzymes CAN change this!

⊖ ΔG – Enzymes CANNOT change this!

The Equilibrium Constant will NEVER be changed by an enzyme

Enzymes

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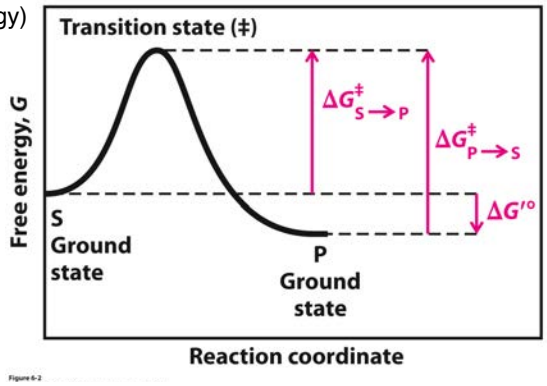


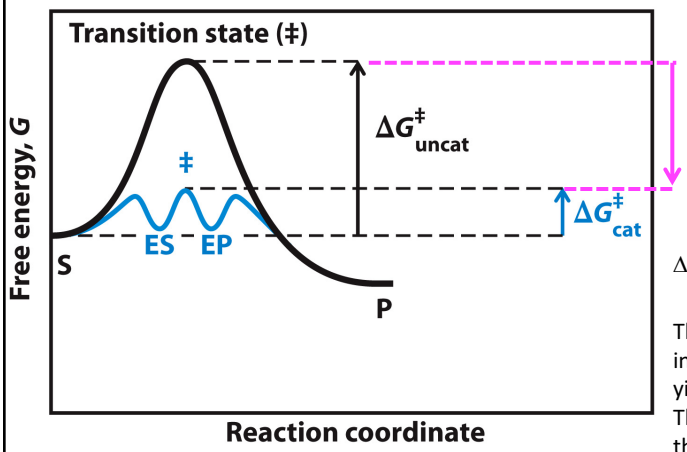
⊕ $\Delta G^{\ddagger}$ – Enzymes CAN change this!

⊖ ΔG – Enzymes CANNOT change this!

The Equilibrium Constant will NEVER be changed by an enzyme

Enzymes

Enzymes Can Decrease $\Delta G^{\ddagger}$



$\Delta\Delta G^{\ddagger}$ ⊖

$\Delta\Delta G^{\ddagger} = \Delta G^{\ddagger}_{cat} - \Delta G^{\ddagger}_{uncat}$

There is always a difference in activation energies that yield a negative $\Delta\Delta G^{\ddagger}$. This is the amount of energy that must be supplied somehow to the reaction by the enzyme

Enzymes

*Catalytic Strategies

versus

Mechanistic Strategies

WHAT must Enzymes do to lower Activation Energies?
-nearly all enzymes do these

HOW do Enzymes lower Activation Energies?
- enzymes may use none, one, or more of these

*Textbook uses this term a bit incorrectly. What they term Catalytic strategies are really those that answer HOW enzymes decrease the activation energy. The HOW-to strategies are really "Mechanistic" strategies.

Enzymes

Catalytic Strategies

- Position Effects: bind substrates where they need to be for reaction (rather than depending on random collisions)
- Polarization of bonds: make substrates more reactive by polarizing bonds (make better nucleophiles, electrophile, or leaving groups) (Electrostatics)
- Strain of bonds: bind substrates in such a way that they "look" like products (put strain on bonds that are to be broken (sessile)) (Geometry)
- De-solvation: assist in removal of water shell around substrates or adding to products upon release (S & P are usually in direct contact with residues at the active site (no water))

Enzymes

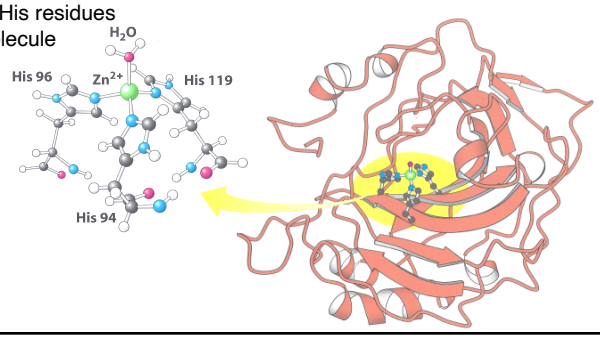
Catalytic Strategies

EXAMPLE: Carbonic Anhydrase

$$CO_2 + H_2O \rightleftharpoons H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$$

Carbonic Acid — Bicarbonate

- Named for the reverse reaction, the dehydration of Carbonic Acid to get water and carbon dioxide
- Mostly β-structure
- It needs a Zn^{2+} metal cofactor
- Held in the active site by 3 His residues
- Fourth ligand is a water molecule



Enzymes

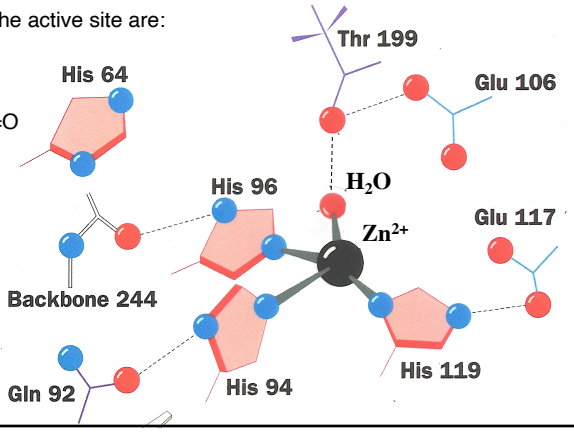
Catalytic Strategies

EXAMPLE: Carbonic Anhydrase

$$CO_2 + H_2O \rightleftharpoons H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$$

Carbonic Acid — Bicarbonate

- Other important residues in the active site are:
  - His-64
  - Thr-199
  - Glu-106
  - Glu-117, Gln-92, 244C=O

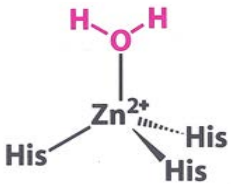


Enzymes

Catalytic Strategies

$CO_2 + H_2O \rightleftharpoons H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$

- Reaction Mechanism

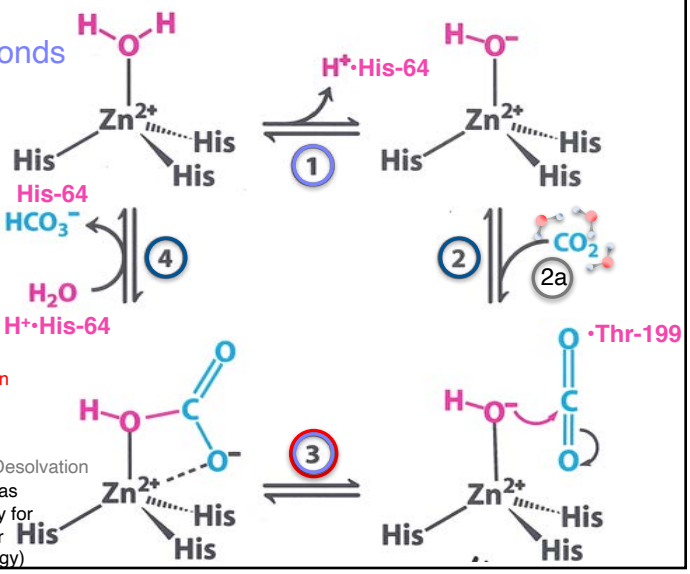


Enzymes

Catalytic Strategies

$CO_2 + H_2O \rightleftharpoons H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$

Position Effects
Polarization of bonds
Strain of bonds
Desolvation



Let's consider how these strategies help "pay" for lowering the activation energy:

Entropy? Position & Strain (but make it more costly)

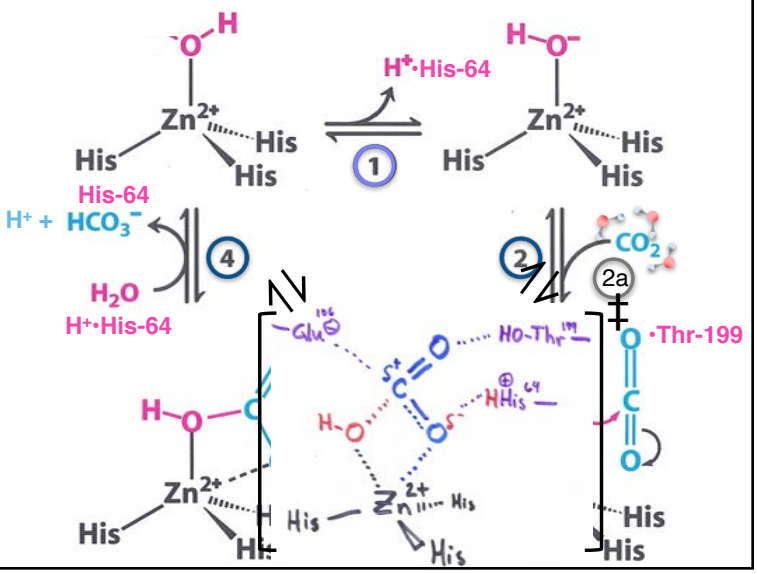
Enthalpy? Polarization & Desolvation (by binding tightly to substrates as they go through the TS, they pay for the loss from entropy AND lower the un-catalyzed activation energy)

Enzymes

Catalytic Strategies

$CO_2 + H_2O \rightleftharpoons H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$

- Transition State Binding

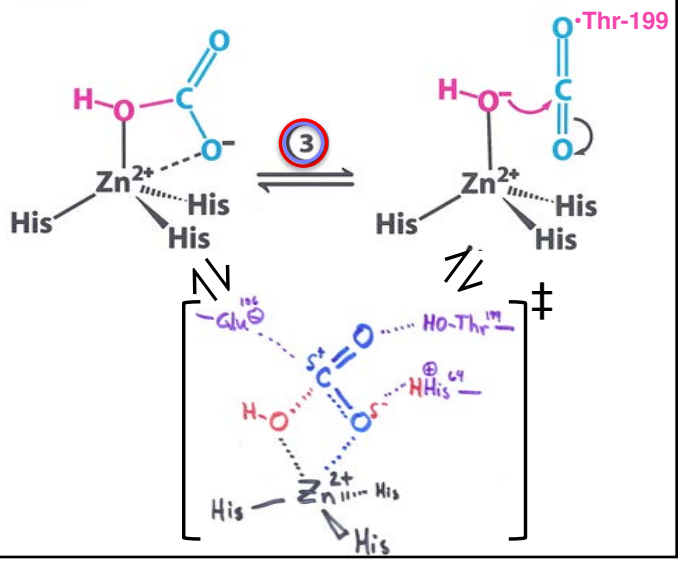


Enzymes

Catalytic Strategies

$CO_2 + H_2O \rightleftharpoons H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$

- Transition State Binding



What is another example of binding the TS?

Enzymes

Illustration of TS Stabilization Idea: Imaginary Stickase

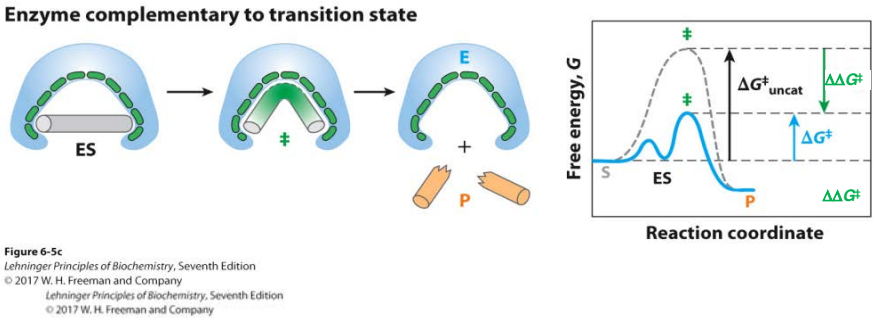


Figure 6-5c
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$\Delta\Delta G^{\ddagger} = \Delta G_M$ (in textbook)

Enzymes

Illustration of TS Stabilization Idea: Imaginary Stickase